1. Determine the Mass of Iron lost by the nail.

17.07 g   Mass of nail before reaction

− 14.42 g   Mass of nail after reaction

2.65 g   mass lost by the nail

= 2.65 g

2. Number of Moles of Iron Used

mass of iron used 2.65 g

Molar mass of Iron = 56

$$Moles = \frac{Mass}{Molar\ mass}$$

$$= \frac{2.65}{56} = 0.04732$$

= 0.04732 moles

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